

TRASFORMAZIONI TERMODINAMICHE

EQUAZIONE DI STATO DEI GAS

$PV=RT$

Pa(N/m²)

m³

J/K

K

P

V

=

R

T

R atmosfera =

287,00

J/K

p atmosfera=

101.325,00

Pa(N/m²)

p₀ p₁

101.325,00 101.325,00

ISOBARA

$R_0T_0/V_0=R_1T_1/V_1$

v₀ v₁

10 20

T₀ T₁

287,16 574,32

$T_0/V_0=T_1/V_1$

p v₀ v₁ T₀ T₁

101.325,00 10 20 287,16 574,32

v₀

10,00

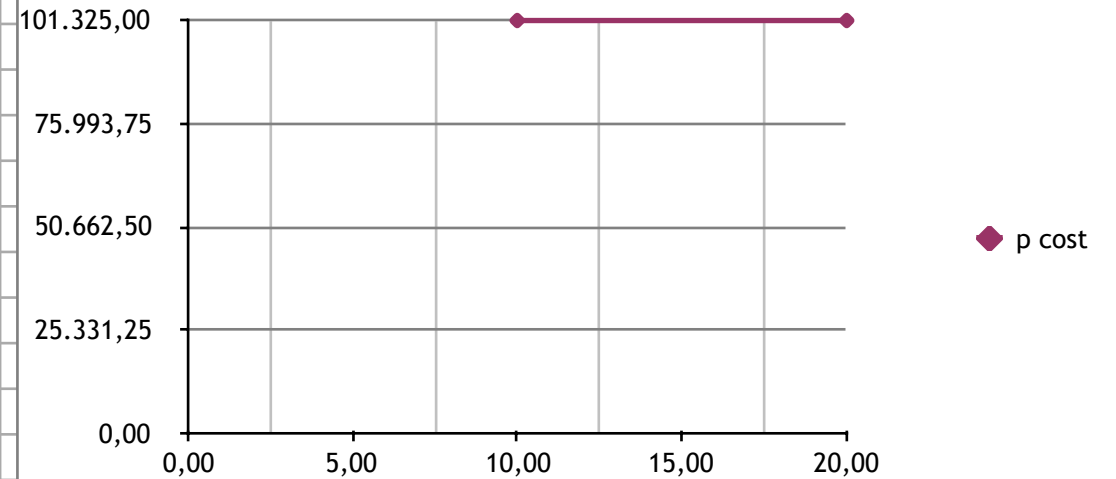
v₁

T₀

T₁

574,32

p cost



ISOCORA

$R_0T_0/P_0=R_1T_1/P_1$

p₀ p₁

101.325,00 151.325,00

v₀ v₁

10 10

T₀ T₁

287,16

v=cost

